Science 8 **Density Calculations Worksheet I ANSWERS** Mr. Birt

Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Class\_\_\_\_\_ Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

density = mass UNITS OF DENSITY

**M**

**D V**

volume solids (g/cm3) liquids (g/mL)

1 kg = 1000 g 1 L = 1000mL 1 g = 1000 mg

1. Find the unknown quantity:

|  |  |  |
| --- | --- | --- |
| a) d = 3 g/mL  V = 100 mL  M = ?  **M = DxV**  **= 3 g/mL x 100 mL**  **= 300 g** | b) d = ?  V = 950 mL  M = 95 g  **D = M**  **V**  **= 95 g**  **950 mL**  **= 0.1 g/mL** | c) d = 0.5 g/cm3  V = ?  M = 20 g  **V = M**  **D**  **= 20 g**  **0.5 g/cm3**  **= 40 cm3** |

2. Find the unknown quantity (CONVERT FIRST to g or mL)

|  |  |  |
| --- | --- | --- |
| a) d = 24 g/mL  V = 1.2 L = **1200** mL  M = ?  **M = DxV**  **= 24 g/mL x 1200 mL**  **= 28,800 g** | b) d = ?  V = 100 mL  M = 1.5 kg = **1500** g  **D = M**  **V**  **= 1500 g**  **100 mL**  **= 15 g/mL** | c) d =  V = 0.52 L = **520** mL  M = 500 mg = **0.5** g  **D = M**  **V**  **= 0.5 g**  **520 mL**  **= 0.001 g/mL** |

**WORD PROBLEMS**

1. A block of aluminum occupies a volume of 15.0 mL and weighs 40.5 g. What is its density?

**D = M = 40.5 g = 40.5 g = 2.7 g/cm3**

**V 15 mL 15 cm3**

2. Mercury metal is poured into a graduated cylinder that holds exactly 22.5 mL. The mercury used to fill the cylinder weighs 306.0 g. From this information, calculate the density of mercury.

**D = M = 306 g = 13.6 g/mL**

**V 22.5 mL**

3. What is the weight of the ethanol that exactly fills a 200.0 mL container?



The density of ethanol is 0.789 g/mL.

**M = DxV = 0.789 g/mL x 200 mL = 157.8 g**

4. A rectangular block of copper metal weighs 1896 g. The dimensions of the block are 8.4 cm by 5.5 cm by 4.6 cm. From this data, what is the density of copper? (hint: find the volume of a block first)

**V = L x W x H = 8.4 cm x 5.5 cm x 4.6 cm = 212.52 cm3**

**D = M = 1896 g = 8.92 g/mL**

**V 212.52 cm3**

5. What volume of silver metal will weigh exactly 2500.0 g. The density of silver is 10.5 g/cm3.

**V = M = 2500 g = 238.1 cm3**

**D 10.5 g/cm3**

6. Find the mass of 250.0 mL of benzene. The density of benzene is 0.8765 g/mL.

**M = DxV = 0.8765 g/mL x 250 mL = 219.13 g**

7. A block of lead has dimensions of 4.50 cm by 5.20 cm by 6.00 cm. The block weighs 1587 g. From this information, calculate the density of lead.

**V = L x W x H = 4.5 cm x 5.2 cm x 6 cm = 140.4 cm3**

**D = M = 1587 g = 11.3 g/cm3**

**V 140.4 cm3**

8. 28.5 g of iron shot is added to a graduated cylinder containing **45.50 mL** of water **(V1)**. The water level rises to the **49.10 mL** **(V2)** mark. From this information, calculate the density of iron.

**V = V2 – V1 = 49.1 – 45.5 = 3.6 mL = 3.6 cm3**

**D = M = 28.5 g = 7.92 g/cm3**

**V 3.6 cm3**